

Cambridge International Examinations Cambridge Ordinary Level

COMBINED SCIENCE

Paper 1 Multiple Choice

Additional Materials: Multiple Choice Answer Sheet Soft clean eraser Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid. Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you. DO **NOT** WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer. Any rough working should be done in this booklet. A copy of the Periodic Table is printed on page 16. Electronic calculators may be used.

This document consists of 15 printed pages and 1 blank page.

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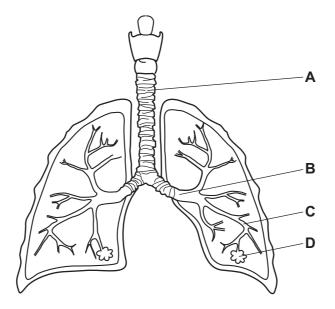
1 hour

May/June 2018



- **1** What is the name of the process by which water passes through a partially permeable membrane?
 - **A** evaporation
 - **B** excretion
 - C osmosis
 - **D** transpiration
- 2 The diagram shows the human breathing system.

Where does diffusion of oxygen and carbon dioxide take place?



3 Four test-tubes contain starch solution and amylase. They are placed in water baths at different temperatures and provided with different pHs, as shown in the table.

After 30 minutes, iodine solution is added to each tube.

In which test-tube do the contents remain yellow-brown?

| | temperature/°C | рН |
|---|----------------|-----|
| Α | 35 | 2.5 |
| в | 35 | 6.9 |
| С | 75 | 2.5 |
| D | 75 | 6.9 |

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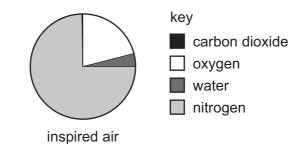
4 A farmer uses faeces and urine from his cattle as fertiliser.

What is the main element provided by fertiliser that the plants use to make proteins?

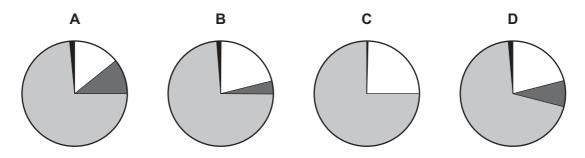
- A carbon dioxide
- B nitrogen
- **C** oxygen
- D water
- 5 What is the name of the process that moves food along the alimentary canal?
 - A absorption
 - B assimilation
 - **C** digestion
 - D peristalsis
- **6** What is transpiration?
 - **A** absorption of water by root hairs
 - **B** loss of water vapour from stomata
 - **C** movement of water up through the xylem
 - **D** wilting
- 7 What is a cause of coronary heart disease?
 - A blockage of the valves in the heart
 - B bursting of the coronary arteries
 - C deposit of fat in the coronary arteries
 - D irregular heartbeat

4

8 The pie chart shows the proportion of gases in inspired air.



Which pie chart represents expired air?



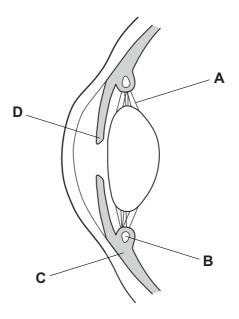
9 The body cannot store amino acids.

Which flow chart correctly shows what happens to excess amino acids in blood?

| Α | excess amino acids in the blood | \rightarrow | broken down in kidney | \rightarrow | urea in the urine | \rightarrow | travel to liver | \rightarrow | urea in the blood |
|---|---------------------------------------|---------------|-----------------------------|---------------|----------------------|---------------|---------------------|---------------|----------------------|
| В | excess amino acids in the blood | \rightarrow | broken down in kidney | \rightarrow | urea in the blood | \rightarrow | travel to liver | \rightarrow | urea in the urine |
| С | excess amino acids in the blood | \rightarrow | broken down in liver | \rightarrow | urea in the urine | \rightarrow | travel to kidney | \rightarrow | urea in the blood |
| D | excess amino acids in the blood | \rightarrow | broken down in liver | \rightarrow | urea in the blood | \rightarrow | travel to kidney | \rightarrow | urea in the urine |

10 The diagram shows a section through part of a human eye.

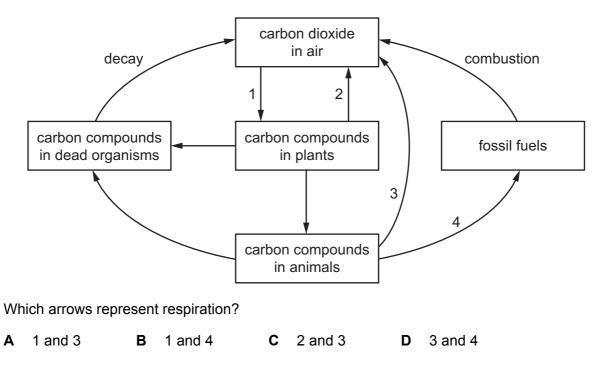
Which structure contains the muscles that contract to control pupil size?



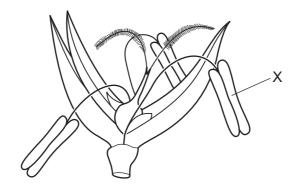
11 Which row best describes some of the effects of alcohol abuse?

| | short-term effect | long-term effect |
|---|----------------------|----------------------|
| Α | addiction | liver disease |
| В | addiction | reduced self-control |
| С | liver disease | addiction |
| D | reduced self-control | liver disease |

12 The diagram shows the carbon cycle.



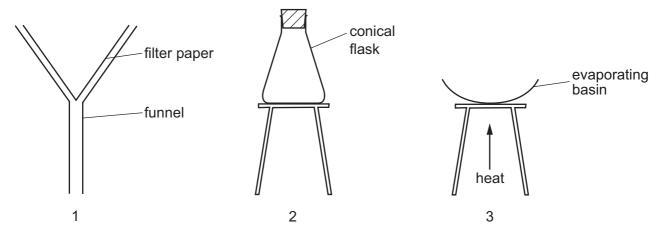
13 The diagram shows a wind pollinated plant.



What is structure X?

- A anther
- B carpel
- **C** petal
- D sepal

14 The diagrams show three sets of apparatus.



Which apparatus is used to obtain separate samples of sand and salt from a mixture of sand and salt solution?

| Α | 1 and 3 | В | 1 only | С | 2 and 3 | D 3 only |
|---|---------|---|--------|---|---------|-----------------|
| | | _ | , | • | | – o o ny |

15 An atom of sodium is represented by $^{23}_{11}$ Na.

What is the number of electrons in this atom?

| | Α | 11 | В | 12 | С | 23 | D | 34 |
|--|---|----|---|----|---|----|---|----|
|--|---|----|---|----|---|----|---|----|

- 16 Which statement about the formation of ions is correct?
 - A Metal atoms gain electrons to form positive ions.
 - B Metal atoms lose electrons to form negative ions.
 - **C** Non-metal atoms gain electrons to form negative ions.
 - D Non-metal atoms lose electrons to form positive ions.
- 17 Which statement about covalent bonding is not correct?
 - **A** A covalent bond forms when a metal atom donates an electron to a non-metal atom.
 - **B** A covalent bond is a pair of shared electrons.
 - **C** The bonding between oxygen and hydrogen is covalent.
 - **D** When atoms form covalent bonds, they get the same electronic configuration as a noble gas.
- **18** The formula of an ammonium ion is NH_4^+ .

The formula of a sulfate ion is SO_4^{2-} .

What is the formula of ammonium sulfate?

A NH_4SO_4 **B** $NH_4(SO_4)_2$ **C** $(NH)_4SO_4$ **D** $(NH_4)_2SO_4$

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- 19 Which balanced equation for the reaction between iron and oxygen is correct?
 - $\textbf{A} \quad Fe_2 \ \textbf{+} \ O_3 \ \rightarrow \ Fe_2O_3$
 - $\textbf{B} \quad 2\text{Fe} \ \textbf{+} \ 3\text{O} \ \rightarrow \ \text{Fe}_2\text{O}_3$
 - $\textbf{C} \quad 4Fe \ + \ 2O_2 \ \rightarrow \ 2Fe_2O_3$
 - $\textbf{D} \quad 4Fe \ \textbf{+} \ 3O_2 \ \rightarrow \ 2Fe_2O_3$
- 20 Which statement about bases is not correct?
 - A Bases dissolved in water turn red litmus blue.
 - **B** Bases neutralise sodium hydroxide solution.
 - **C** Bases react with acids to form salts.
 - **D** Bases react with ammonium salts to form ammonia.
- **21** P, Q, R and S are four elements.

The letters are not their chemical symbols.

| element | physical state at room temperature | number of electrons in outer shell | metal or non-metal |
|---------|------------------------------------|---------------------------------------|-----------------------|
| Р | gas | 2,6 | non-metal |
| Q | gas | 2,7 | non-metal |
| R | solid | 2,8,2 | metal |
| S | gas | 2,8,7 | non-metal |

Which elements are in the same group of the Periodic Table?

 A
 P and Q
 B
 P and S
 C
 Q and S
 D
 R and S

22 A metal is used to make a pipe to transport hydrochloric acid.

Which metal is suitable for making the pipe?

- A copper
- **B** iron
- C magnesium
- D zinc

23 The table shows some metals and their uses.

For which metal is the correct reason given for the stated use?

| | metal | use | reason |
|---|-----------|--------------------------------|---------------------------------|
| Α | aluminium | manufacture of aeroplane wings | strength and high density |
| в | copper | electrical wiring | good conductor of heat |
| С | iron | manufacturing stainless steel | rusts |
| D | zinc | galvanising iron | zinc is more reactive than iron |

24 A sample of polluted air is shaken with 50 cm³ of distilled water and the pH of the resulting solution is measured.

The experiment is repeated with the same volume of unpolluted air.

The results are shown.

| sample | рН |
|----------------|----|
| unpolluted air | 6 |
| polluted air | 2 |

Which statement explains the pH of the polluted air?

- **A** It is polluted with carbon dioxide.
- **B** It is polluted with carbon monoxide.
- **C** It is polluted with lead compounds.
- **D** It is polluted with sulfur dioxide.
- 25 Which substance produces hydrogen gas when it reacts with dilute hydrochloric acid?
 - A magnesium
 - **B** magnesium carbonate
 - **C** magnesium hydroxide
 - **D** magnesium oxide
- 26 Which molecular formula represents an alkane?

| A C_2H_2 B C_3H_8 C C_4H_8 | C_3H_8 | C C₄H ₈ | D C_5H_{10} |
|---|----------|---------------------------|----------------------|
|---|----------|---------------------------|----------------------|

27 A reaction of ethanol is shown.

 $\mathsf{CH}_3\mathsf{CH}_2\mathsf{OH}~+~3\mathsf{O}_2~\rightarrow~2\mathsf{CO}_2~+~3\mathsf{H}_2\mathsf{O}$

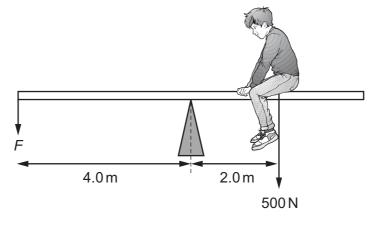
Which statement about this reaction is not correct?

- **A** One of the products turns lime-water cloudy.
- **B** The ethanol is a fuel.
- **C** The ethanol is being reduced.
- **D** The reaction is exothermic.
- **28** The gradient of the line on a graph gives the acceleration of a moving object.

What are the quantities on the horizontal and vertical axes of this graph?

| | quantity on horizontal axis | quantity on vertical axis |
|---|--------------------------------|---------------------------|
| Α | speed | distance |
| в | speed | time |
| С | time | distance |
| D | time | speed |

- 29 Which statement concerning the mass of a body is incorrect?
 - **A** Mass can be measured using an appropriate balance.
 - **B** Mass experiences a force due to gravitational attraction.
 - **C** Mass is a measure of the amount of substance in a body.
 - **D** Mass is varied by changes in the strength of a gravitational field.



30 The diagram shows a boy of weight 500 N sitting on a see-saw. He sits 2.0 m from the pivot.

What is the force F needed to balance the see-saw?

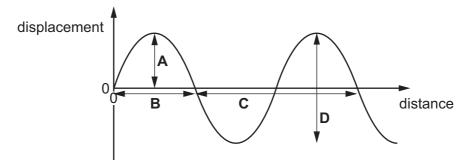
A 250 N **B** 750 N **C** 1000 N **D** 3000 N

31 How much work is done in lifting a mass of 70 g vertically through a distance of 6 m? (gravitational field strength is 10 N/kg.)

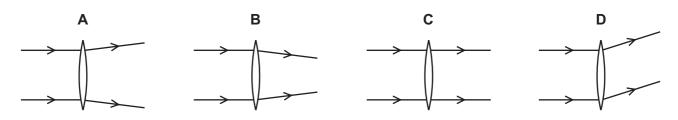
A 0.42 J **B** 4.2 J **C** 420 J **D** 4200 J

- **32** What makes the metal mercury a suitable liquid for use in a thermometer?
 - A It expands uniformly when heated.
 - **B** It is a poor conductor of heat.
 - **C** It is more dense than glass.
 - **D** It reacts slowly to changes in a temperature.
- **33** The diagram shows the displacement of the particles in a wave.

Which value is multiplied by the frequency to give the speed of the wave?



34 Which diagram correctly shows the path of two rays of light after they pass through a thin converging lens?



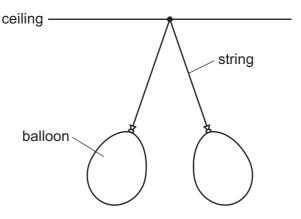
35 The diagram shows the main components of the electromagnetic spectrum.

| P X-rays Q | visible light infra-red | R radio waves |
|------------|-------------------------|---------------|
|------------|-------------------------|---------------|

What are the components P, Q and R?

| | Р | Q | R |
|---|------------|--------------|--------------|
| Α | gamma-rays | microwaves | ultra-violet |
| в | gamma-rays | ultra-violet | microwaves |
| С | microwaves | gamma-rays | ultra-violet |
| D | microwaves | ultra-violet | gamma-rays |

36 Two balloons are suspended from the ceiling by string and have moved apart as shown.

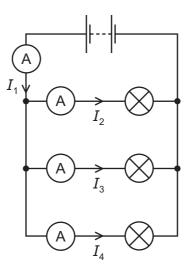


Which statement is correct?

- **A** One is charged and the other is uncharged.
- **B** They are uncharged.
- C They have like charges.
- **D** They have unlike charges.

12

37 A student sets up the circuit shown.



The currents measured by the ammeters are shown.

Which equation is correct?

- **A** $I_1 = I_2 + I_3 + I_4$
- **B** $I_1 = I_2 = I_3 = I_4$
- **C** $I_2 + I_3 = I_4 + I_1$
- **D** $I_4 = I_3 + I_2 + I_1$
- **38** A 5W electric night light is used for 8 hours per day over a period of 7 days.

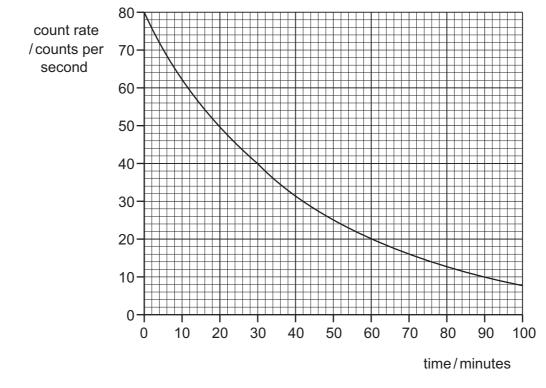
How much electrical energy is transferred to the night light?

A 280 J **B** 16800 J **C** 144000 J **D** 1008000 J

39 In a simple a.c. generator, a coil is rotated in a uniform magnetic field.

Which action would **not** increase the size of the maximum e.m.f. generated?

- A increasing the number of turns of the coil
- **B** increasing the rate of rotation of the coil
- **C** increasing the resistance of the coil
- **D** increasing the strength of the magnetic field



40 The graph shows how the count rate measured from a radioactive nuclide changes with time.

What is the half-life of this nuclide?

| A 17 minutes B 25 minutes C 30 minutes D 50 minutes | Α | 17 minutes | В | 25 minutes | С | 30 minutes | D | 50 minutes | |
|---|---|------------|---|------------|---|------------|---|------------|--|
|---|---|------------|---|------------|---|------------|---|------------|--|

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15

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The volume of one mole of any gas is $24\,dm^3$ at room temperature and pressure (r.t.p.).

Lu Iutetium 175 103 Lr lawrencium

Yby 173 173 173 173 102 NO

Tm 169 101 Md -

Er 167 167 100 100 -

holmium 165 99 einsteinium

Dy dysprosium 163 98 98 Cf

Tb 159 97 97 Bk berkelium

Gd addolinium 157 96 Cm cunium cunium

Eu ^{europium} 152 95 Am amenicium

Samarium 150 94 94 Pu Pu Putonium

promethium 33 93 93 - - hium - - neptunium

144 02 92 92 02 238 238

Praseodymium 141 91 93 Pa protactinium 231

Cerium 140 90 90 Th Thorium

La lanthanum 139 89 89 actinium

actinoids

lanthanoids

The Periodic Table of Elements

| | | _ | |
|--------|------|----|--|
| © UCLE | S 20 | 18 | |

| | NIII | He ² | helium 4 | 10 N | neon 20 | 18 | Ar | argon 40 | 36 | Ϋ́ | krypton 84 | 54 | Xe | xenon 131 | 86 | Rn | radon _ | | | |
|-------|------|-----------------|---------------|--------------|------------------------------|----|----|------------------|----|----|-----------------|----|--------|------------------|-------|-------------|-----------------|--------|-----------|--------------------|
| | ٨I | | | бЦ | fluorine 19 | 17 | Cl | chlorine 35.5 | 35 | Ъ | bromine 80 | 53 | Ι | iodine 127 | 85 | At | astatine - | | | |
| | N | | | ∞ C | oxygen 16 | 16 | ა | sulfur 32 | 34 | Se | selenium 79 | 52 | Te | tellurium 128 | 84 | Ро | polonium – | 116 | 2 | livermorium – |
| | > | | | ۲ N | nitrogen 14 | 15 | ۵. | phosphorus 31 | 33 | As | arsenic 75 | 51 | Sb | antimony 122 | 83 | Bi | bismuth 209 | | | |
| | ≥ | | | ه ر | carbon 12 | 14 | Si | silicon 28 | 32 | Ge | germanium 73 | 50 | Sn | tin 119 | 82 | Pb | lead 207 | 114 | Γl | flerovium - |
| | ≡ | | | 5 U |) boron 1 | 13 | Ρl | aluminium 27 | 31 | Ga | gallium 70 | 49 | In | indium 115 | 81 | 11 | thallium 204 | | | |
| | | | | | | | | | 30 | Zn | zinc 65 | 48 | Cd | cadmium 112 | 80 | Hg | mercury 201 | 112 | Cn | copernicium - |
| | | | | | | | | | 29 | Cu | copper 64 | 47 | Ag | silver 108 | 79 | Au | gold 197 | 111 | Rg | roentgenium - |
| dn | | | | | | | | | 28 | ïZ | nickel 59 | 46 | Pd | palladium 106 | 78 | Ъ | platinum 195 | 110 | Ds | darmstadtium - |
| Group | | | | | | | | | 27 | ပိ | cobalt 59 | 45 | Rh | rhodium 103 | 77 | Ir | iridium 192 | 109 | Mt | meitnerium - |
| | | - T | hydrogen 1 | | | | | | 26 | Ее | iron 56 | 44 | Ru | ruthenium 101 | 76 | SO | osmium 190 | 108 | Hs | hassium – |
| | | | | | | _ | | | 25 | Mn | manganese 55 | 43 | Ц | technetium - | 75 | Re | rhenium 186 | 107 | Bh | bohrium – |
| | | | | - lod | ass | | | | 24 | ŗ | chromium 52 | 42 | Mo | molybdenum 96 | 74 | ≥ | tungsten 184 | 106 | Sg | seaborgium - |
| | | | Key | atomic symbo | name relative atomic mass | | | | 23 | > | vanadium 51 | 41 | qN | niobium 93 | 73 | Та | tantalum 181 | 105 | Db | dubnium – |
| | | | | atc | | | | | 22 | F | titanium 48 | 40 | Zr | zirconium 91 | 72 | Ŧ | hafnium 178 | 104 | Ŗ | rutherfordium - |
| | | | | | | 1 | | | 21 | Sc | scandium 45 | 39 | ≻ | yttrium 89 | 57-71 | lanthanoids | | 89-103 | actinoids | |
| | = | | | 4 U | beryllium 9 | 12 | Mg | magnesium 24 | 20 | Ca | calcium 40 | 38 | ي د | strontium 88 | 56 | Ba | barium 137 | 88 | Ra | radium - |
| | _ | | | 3 | lithium 7 | 11 | Na | sodium 23 | 19 | ¥ | potassium 39 | 37 | Rb | rubidium 85 | 55 | Cs | caesium 133 | 87 | ч | francium - |

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